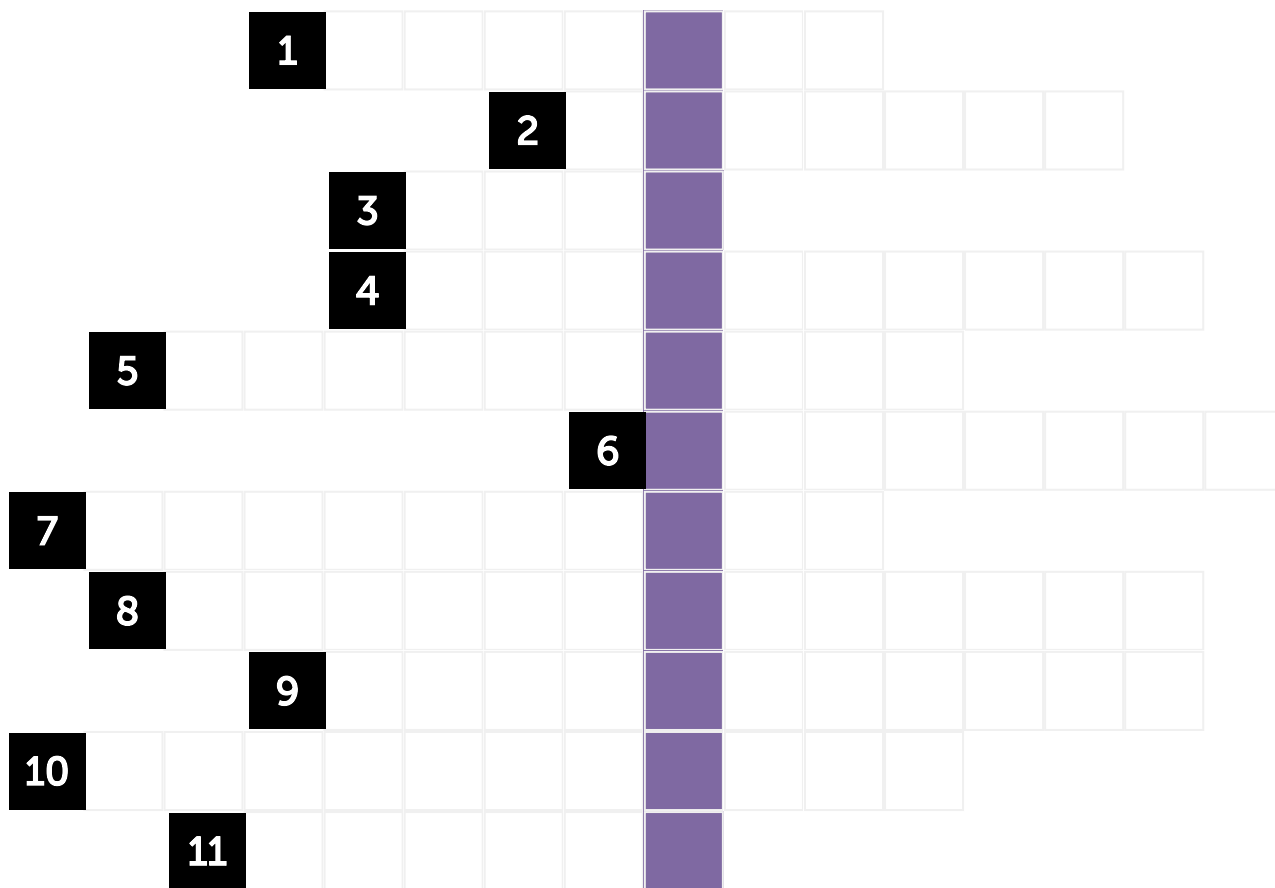


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F	Y	C	L	N	A	G	O	S	B	Y	O	Q	K	R
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H	C	H	Q	R	T	T	D	D	O	Z	C	W	A	Q
D	T	Q	K	G	R	S	E	P	L	O	E	S	E	N
Z	I	T	T	Y	M	Y	R	A	L	U	N	I	R	Z
K	V	J	F	G	P	L	U	C	I	F	T	K	A	X
V	A	H	K	R	R	A	T	J	S	H	R	R	E	N
O	T	G	N	A	E	T	A	W	I	I	A	I	C	F
L	I	A	N	T	S	A	R	L	O	H	T	K	A	I
U	O	L	P	E	S	C	E	T	N	N	I	D	F	C
M	N	A	F	L	U	H	P	N	S	L	O	T	R	I
E	K	A	K	D	R	F	M	P	I	I	N	T	U	B
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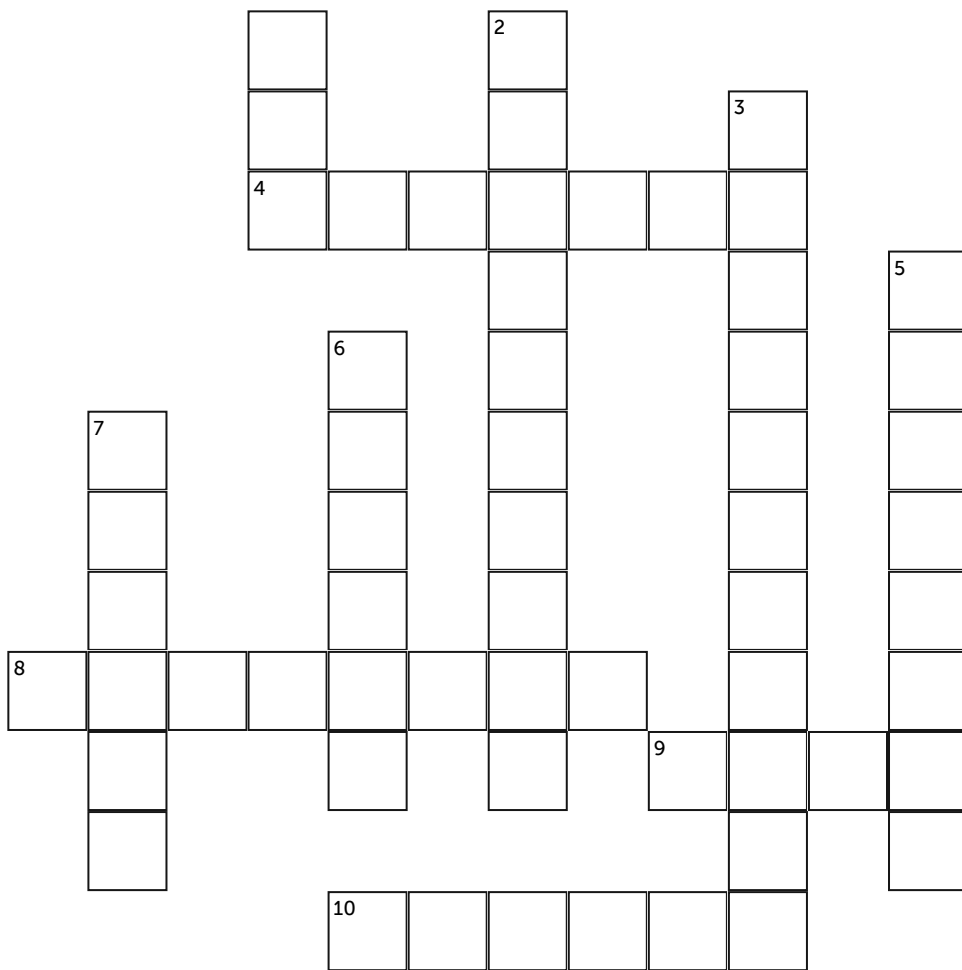
HINTS

surfacearea catalyst temperature concentration rate volume pressure
 collisions activation energy



HINTS

- 1 A word to describe the gradient of a faster rate
- 2 You can either calculate rate of reaction using change in reactant or _____
- 3 _____ of reaction
- 4 For a reaction to happen, collisions must be this
- 5 Enzymes are _____ catalysts
- 6 A factor that affects only reactions with gases
- 7 The type of energy needed to start a reaction
- 8 A factor that affects only reactions with liquids
- 9 A factor that only affects reactions with solids
- 10 The pathway provided during a reaction with a catalyst
- 11 Doesn't affect the rate of reaction



1	2	3	4	5	6	7	8	9	10
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ACROSS

- 4 Increasing the surface area means making the particles _____
- 8 Collisions become more _____ when the rate increases
- 9 Rate of reaction can be monitored by watching the change in _____
- 10 Rate of reaction can be monitored by watching the change in _____

DOWN

- 1 Pressure only affects reactions with this
- 2 These need to be successful
- 3 When an insoluble solid is formed
- 5 Lowers the activation energy
- 6 Rate of reaction can be monitored by watching the change in _____
- 7 Increasing the temperature provides more _____