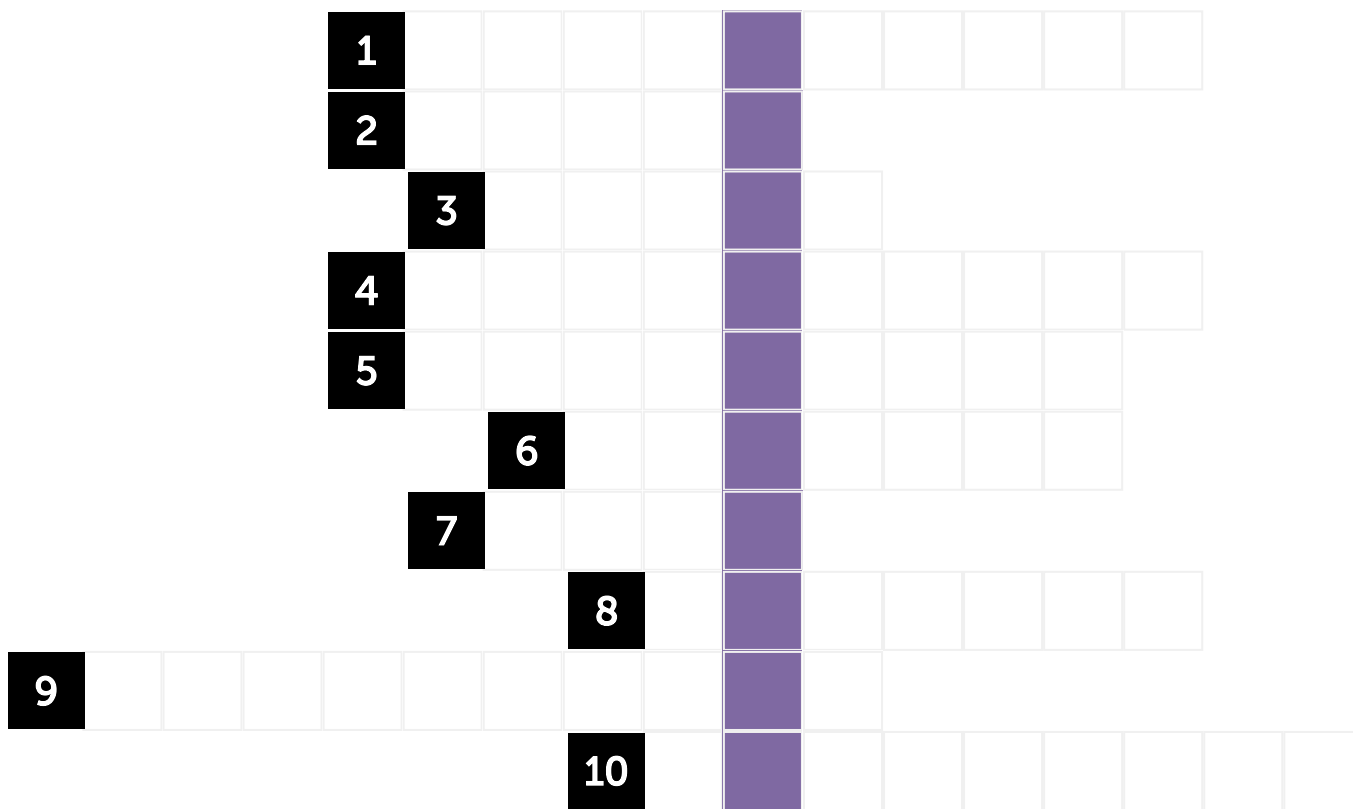


N	K	Z	D	C	U	Q	A	F	P	S	B	E	I	M
V	H	W	I	B	Q	D	E	E	Z	H	K	L	N	A
V	X	O	S	N	Q	V	P	V	F	L	N	E	U	N
A	Y	N	P	O	K	H	M	O	U	H	O	C	G	L
P	L	N	L	I	F	T	M	B	D	V	I	T	E	E
N	S	U	A	T	H	M	L	N	Z	S	T	R	T	C
O	T	C	C	C	D	M	P	P	J	Z	C	O	R	F
I	N	L	E	A	T	O	U	S	Z	C	U	N	E	T
T	E	E	M	E	H	L	O	T	W	U	D	O	N	O
A	M	U	E	R	H	E	R	Q	F	J	E	Q	I	L
D	E	S	N	J	Q	C	G	H	D	U	R	H	S	R
I	L	P	T	S	B	U	Z	P	P	J	M	T	F	T
X	E	N	H	Q	E	L	V	I	C	I	B	F	Z	M
O	H	Y	P	C	H	E	B	B	Q	V	V	I	O	P
H	Q	R	B	K	Y	U	D	Z	M	Y	M	W	J	Y

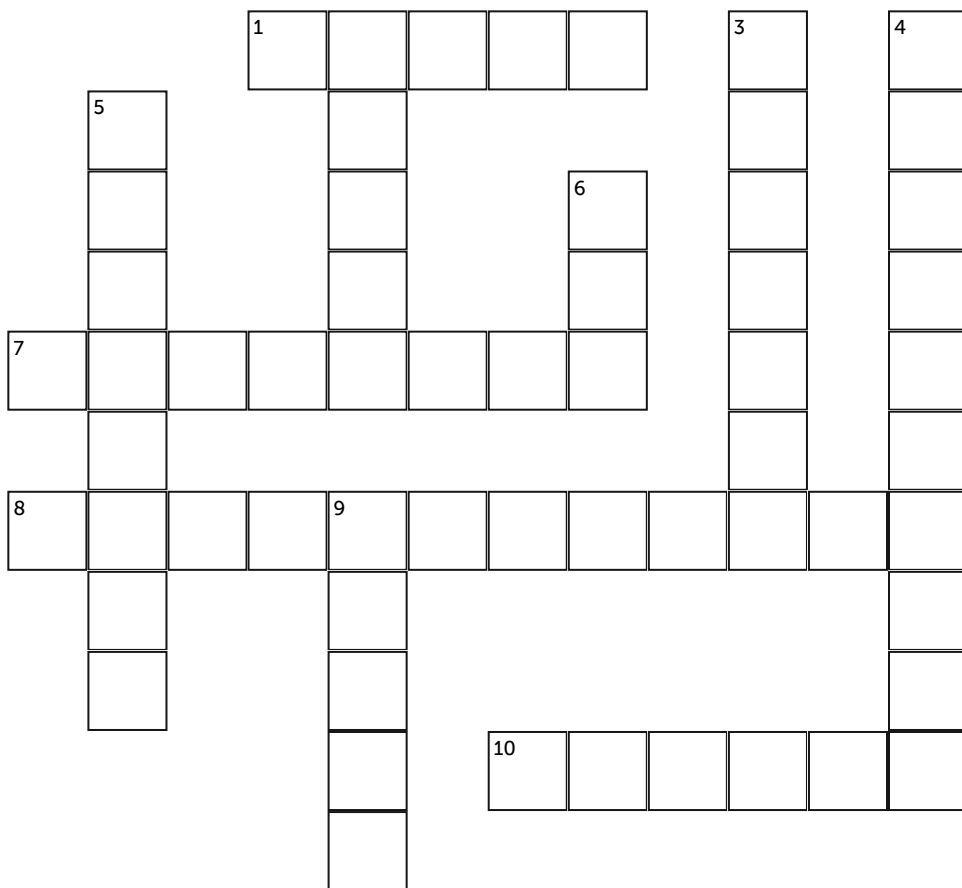
### HINTS

electron    oxidation    reduction    nucleus    group    element    displacement    reaction  
inert    molecule



### HINTS

- 1 The colours of Noble gases
- 2 The colour of bromine at room temperature
- 3 The reason group 0 elements don't react
- 4 One of the reasons group 0 elements have many uses
- 5 The loss of electrons
- 6 When the electrons of group one elements are further away from this, the metal becomes more reactive
- 7 Alkali metals are unusually \_\_\_\_\_
- 8 When the group 1 metals react with water, they will all start \_\_\_\_\_
- 9 The force experienced between the nucleus and electrons of an atom
- 10 The gain of electrons



1	2	3	4	5	6	7
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### ACROSS

- 1 The number of electrons on the outershell of Halogens
- 7 The charge of an atom's nucleus
- 8 When a more reactive element will replace a less reactive element in a compound
- 10 The name given to halogen ions

### DOWN

- 2 The number of electrons on the outershell of Noble Gases
- 3 Alkali metals are stored in oil, as they will \_\_\_\_\_ in air
- 4 What the term 'inert' means
- 5 The most reactive halogen
- 6 The number of electrons on the outershell of Alkali Metals
- 9 The colour of the flame produced by potassium reacting with water